



Nickel-nitride composite: An eco-friendly and efficient alternative to platinum for electrocatalytic hydrogen production

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ABSTRACT

Water electrolysis using the hydrogen evolution reaction (HER) is a promising method for sustainable hydrogen production. Platinum-based catalysts have traditionally been the most efficient HER catalysts, but their scarcity and sluggish water dissociation limit their practical applications. Here we report on a novel superhydrophilic catalyst of nickel supported on nickel molybdenum nitride (Ni/Ni_{0.8}Mo_{4.2}N₆) that outperforms platinum-based nanomaterials. Despite the low catalytic activity of Ni or Ni_{0.8}Mo_{4.2}N₆ alone, their optimized composite exhibits exceptional HER activity, with respective 500% and 150% increases in the exchange current and turnover frequency compared to commercial Pt/C. Density of states calculations reveal a decrease in electron density of the supported nickel in Ni/Ni_{0.8}Mo_{4.2}N₆, leading to a lower free energy of the HER. These findings demonstrate a powerful electron-engineering strategy for designing supported electrocatalysts with outstanding performance for the HER and related processes.

1. Introduction

Hydrogen is increasingly regarded as a versatile and eco-friendly energy source to help achieve net-zero targets for CO₂ emissions [1,2]. Electrochemical water-splitting using renewable electricity from sources such as solar and wind power can offer 'green' hydrogen with no CO₂ emissions [3,4]. The efficiency of such technology depends on the selectivity, intrinsic activity, and stability of the electrocatalysts used for the water splitting reactions. Pt is currently the dominant electrocatalyst for cathodic HER due to its almost perfect Gibbs free energy for H⁺ adsorption ($\Delta G_H \approx 0$) [5–8]. However, the scarcity [9,10] and resulting high costs of Pt limit its sustainable use for practical applications, and in industrially-relevant alkaline conditions the activity is hindered by a sluggish water dissociation step [11–13] leading to low efficiencies in

both water-alkali and chlor-alkali electrolyzers [2].

As alternatives to Pt, nickel (Ni) based electrocatalysts are known to efficiently cleave H-OH bonds [14] and find use as industrial HER catalysts in basic solutions [15]. However, these materials have shortcomings as intrinsically hydrophobic surfaces make infiltration of aqueous electrolyte difficult [16], and strong adsorption of H⁺ at active sites impedes the progress of HER [5,17]. Various strategies such as strain engineering [18–20], electronic regulation [21–23], and hydrogen spillover effects [24–27] have been employed to optimize the properties of nickel based catalysts, but their activity has remained inferior to that of Pt [8]. Other active HER catalysts such as metal nitrides with excellent corrosion resistance, electrical conductivity and Pt-like electronic structures, have also been discovered [28,29], but to date no catalyst that outperforms Pt in terms of overpotential and

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current density has been reported.

We have explored two chemical strategies for improving the performance of Mo_5N_6 which was previously reported as a promising HER catalyst [28]; Ni-doping and introducing supported Ni nanoparticles. Ni doping has previously been used to improve properties in MoN_x [30,31]. Supported Ni nanoparticles on metal nitrides (such as MoN [32] and Ni_3N [33]) are also exhibit good HER activity. A combination of both strategies is found to give an outstanding $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite catalyst which exhibits a superhydrophilic surface with nickel nanoparticles uniformly dispersed on $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ nanosheets. Electron transfer from Ni to $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ improves the water activation and weakens the adsorption of H^* on Ni, and hence dramatically enhances the activity of Ni towards HER. As a precious-metal-free material for HER in alkaline solution, $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ exhibits a small overpotential (20 mV at 10 mA cm^{-2}), high turnover frequency ($\text{TOF} = 0.29 \text{ s}^{-1}$ at 50 mV overpotential) and exchange current density ($J_0 = 2.89 \text{ mA cm}^{-2}$), and robust stability (only 26 mV deactivation after 1000 h operation), outperforming the benchmark commercial Pt/C catalyst.

2. Experimental section

2.1. Synthesis of materials

In a typical preparation of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ catalyst by hydrothermal method, 0.746 g nickel acetate ($\text{Ni}(\text{CH}_3\text{COO})_2 \cdot 4 \text{ H}_2\text{O}$) and 0.241 g sodium molybdate dihydrate ($\text{Na}_2\text{MoO}_4 \cdot 2 \text{ H}_2\text{O}$) are dispersed into 50 mL water/ethylene glycol ($\text{Ni}:\text{Mo}_{\text{molar}} \text{ ratio} = 3:1$; water:ethylene glycol_{volume} ratio = 1:1). After ultrasonic dissolution, the solution is poured into the autoclave and the temperature and time are set to 120°C and 12 h, respectively, for the synthesis of Ni-Mo precursors (Fig. S1a). After cooling to room temperature, wash repeatedly with ethanol and water until the filtrate becomes colorless. The resulting sample is then dried sufficiently. The precursor is calcined in a tube furnace under a flowing NH_3 atmosphere. The temperature is raised at a rate of 5°C min^{-1} until it reaches 500°C and then maintained for 4 h. The obtained black product is labeled as $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ (Fig. S1b). For comparison, the $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is prepared by treating $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ in sulfuric acid (0.5 M) for 2 h. Metallic Ni can also be synthesized through a hydrothermal reaction in the absence of $\text{Na}_2\text{MoO}_4 \cdot 2 \text{ H}_2\text{O}$, by flowing a mixture of H_2/Ar (5% H_2) over the reactants at a temperature of 500°C for 2 h, with a heating rate of 5°C min^{-1} . All the reagents were purchased from Sinopharm Chemical Reagent in analytical grade and used as received without further purification.

2.2. Materials characterization

Powder X-ray diffraction measurements were carried out using the Rigaku Miniflex 600 instrument with Cu K α radiation ($\lambda = 1.54178 \text{ \AA}$). The scanning range was $10\text{--}80^\circ$ at 1° min^{-1} . Morphologies and microstructures were characterized by a scanning electron microscope (SEM, Hitachi S-4800) and transmission electron microscope (TEM, FEI Tecnai F20). X-ray photoelectron spectroscopy (XPS) was used to analyze surface valence with a VG ESCALAB MKII instrument and Mg K α excitation source. For the XPS analysis, carbon was used as an internal standard with a binding energy of C1s = 284.8 eV. The content of the metal in electrolyte before and after stability are tested by the inductively coupled plasma optical emission spectrometry (ICP-OES) on SPECTRO ARCOS. The content of oxygen and nitrogen are tested by oxygen and nitrogen analyzer (HORIBA EMGA-620 W). The contact angle is measured by Data physics scientific OCA25 Surface Analyser at room temperature. XANES (X-ray absorption near-edge structure) spectra for Ni K-edge were obtained at BL14W1 beamline in Shanghai Synchrotron Radiation Facility (SSRF) using transmission mode. The spectra were collected in a range spanning approximately 200 eV below to 800 eV above the edge, with a step size of 0.5 eV in the near-edge region. The dwell time for each step was set to 1 s. At the Spallation Neutron Source

of Oak Ridge National Laboratory, data on powder neutron diffraction were gathered using the NOMAD beamline [34]. Further details regarding the experimental and data processing procedures can be found in the Supporting Information.

2.3. Electrochemical tests

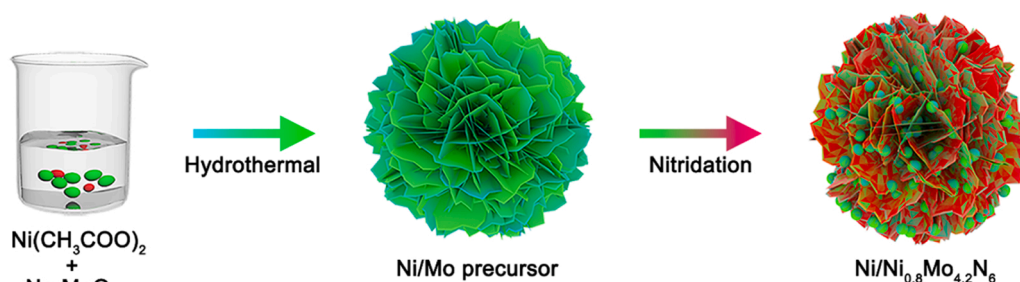
For the experiment, a graphite rod served as the counter electrode while Ag/AgCl (saturated KCl aqueous solution) was used as the reference electrode. To prepare the working electrode, a catalyst ink was created by dispersing the catalyst in a 1:1 mixture of water and isopropyl alcohol with a concentration of 8.0 mg mL^{-1} . A small amount (10 μL , 5.0 wt%) of Nafion is added to serve as a binder. After sonication for 30 min to ensure a uniform mixture, 10.0 μL of the ink is applied to a polished glassy carbon electrode (diameter: 5.0 mm). More information about testing the performance of all electrocatalysts and calibrating the reference electrode in the HER experiment is available in the Supporting Information.

3. Results and Discussion

The $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite and separate $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and Ni nanoparticle samples were generated through hydrothermal chemistry as described in Scheme 1. Neutron diffraction (Fig. 1a) demonstrates that the former sample is a composite of Ni and $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. The structure refinement results show that Ni has been substituted into the hexagonal Mo_5N_6 structure, with the refined composition of $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ (space group, $P6_3/m$; $a = 4.9443(8)$, $c = 11.1280(2) \text{ \AA}$), which comprises alternate layers of MoN_6 triangular prisms and octahedra. Ni partially occupies an octahedral site but shifts towards a trigonal prismatic site (Fig. 1b). Powder X-ray diffraction (XRD) in Fig. 1c shows that cubic Ni and hexagonal $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ phases are both observed for the $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite (as compared to patterns for the individual components in Fig. S2). Rietveld XRD analysis shows that the content of cubic nickel is 48.6 wt%, in agreement with ICP-OES results (Table S1). The Ni particle size in $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ calculated from the Scherrer formula is 8.0 nm, smaller than that for the pristine Ni sample (30 nm). Scanning electron microscope (SEM) shows the $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite to have a marigold-like morphology (Fig. S1).

The overlapping crystal lattices (Fig. 1d-g) suggests that the nanodomains of supported Ni and $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ are contiguous. High-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) and elemental mapping shows the highly dispersed nature of the Ni nanoparticles (NiPs) on the composite surface (Figs. 1h-l and S3-S4). The $\sim 8.0 \text{ nm}$ NiPs domains visible in Fig. 1l evidence a high specific surface area (S_{BET}) of $53 \text{ m}^2 \text{ g}^{-1}$, larger than that of the pristine Ni ($23 \text{ m}^2 \text{ g}^{-1}$). The STEM images in Fig. S5 reveal that homogeneously dispersed salient NiPs on $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ disappear during acidic etching with the formation of mesopore for $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. The diameter of the pores is 8 nm (Fig. S6 and Table S2), which is exactly the same as the size of the NiPs.

Compared to $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$, the negative shift in binding energy for the Mo 3d core level in Fig. 2a suggests an electron rich state of Mo in $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ [35]. Meanwhile, the electron transfers from Ni to the adjacent $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ (suggested by the positive binding energy shift observed for the Ni 2p core level relative to metallic Ni in Fig. 2b) leads to a loss of electron charge density on the Ni. Moreover, compared with $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$, the oxidation peak of Ni 2p in $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ shows a partial positive shift, which is consistent with Mo 3d XPS. Fig. 2c shows the Ni K-edge X-ray adsorption near-edge structure (XANES) spectra of the samples compared to Ni foil. The absorption edge of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ locates very close to that of the Ni foil indicates that Ni in $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ resembles its metallic state [17]. However, the enhanced white line intensity and slight shift toward higher energy of the absorption edge (inset Fig. 2c) suggest that Ni in $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ has a partial positive charge [36]. The confirmed electron-poor Ni supported on



Scheme 1. Schematic diagram of the synthesis for the $\text{Ni/Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite.

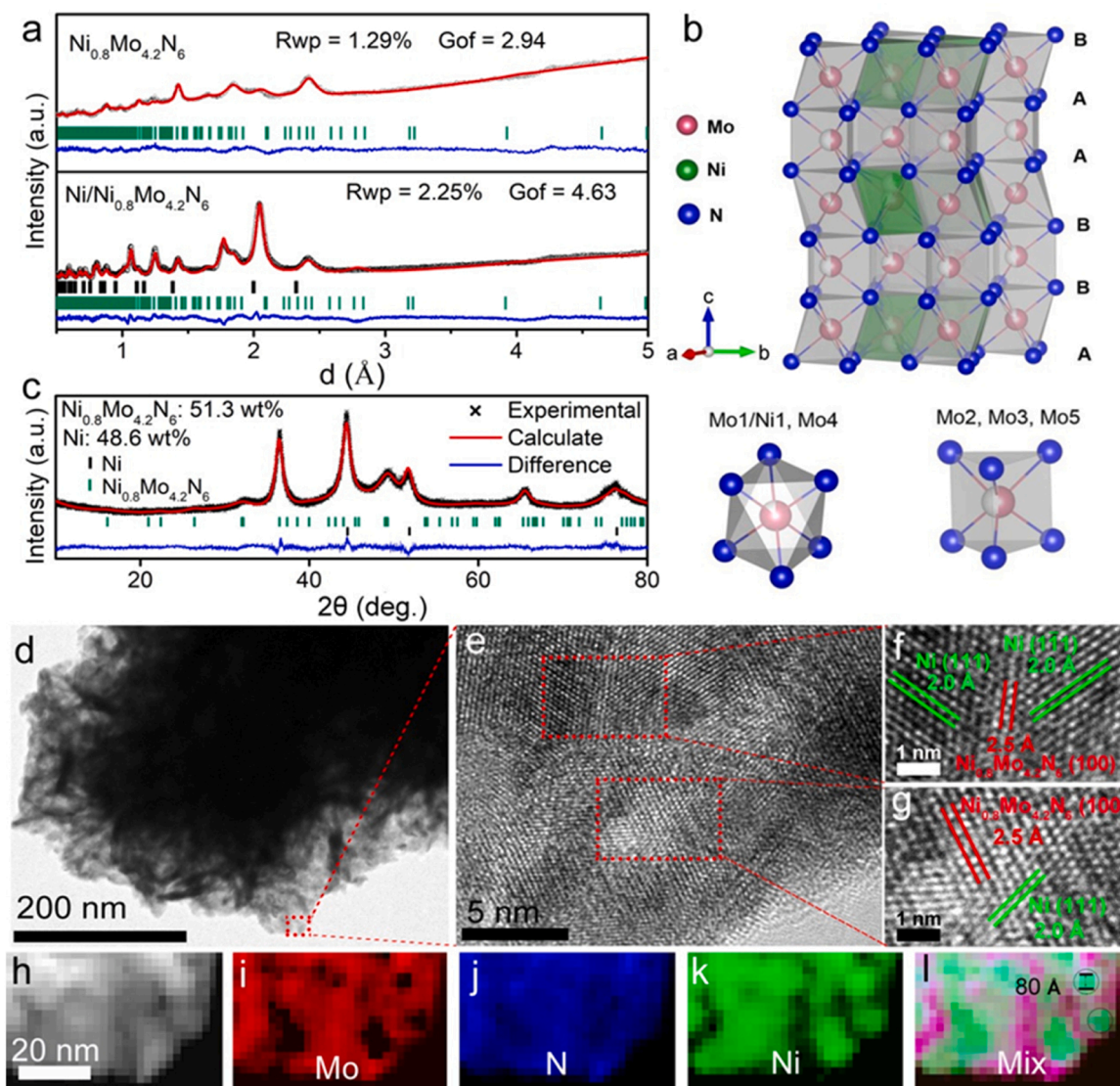


Fig. 1. (a) Neutron diffraction patterns and Rietveld refinement of $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and $\text{Ni/Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. (b) Crystal structure of $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. (c) The Rietveld refined XRD ($\lambda = 1.54178 \text{ \AA}$, Cu K α) pattern of $\text{Ni/Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. (d-g) The TEM and HR-TEM images. (h-l) HAADF-STEM image and the energy dispersive X-ray spectroscopy (EDS) element mapping for Ni, Mo, N element.

$\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is important in such systems to weaken H^* adsorption and thus boost the HER [21,22,37].

As HER is a gas–liquid–solid three-phase process, a superhydrophilic surface is favorable to promote the infiltration of aqueous electrolyte and facilitate mass transfer [16,38,39]. In contrast to the Pt/C (20 wt%) catalyst which is hydrophobic, liquid-solid contact-angle measurements in Fig. 2d and supplementary videos reveal that $\text{Ni/Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is

superhydrophilic. The dispersion of Ni nanostructures on the surface of the nitride, as well as the highly porous marigold structure, aids the structural superhydrophilicity, beyond that of the nitride or metallic Ni alone (Fig. S7).

HER activity of the $\text{Ni/Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ composite in comparison to the individual Ni and $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ components and a benchmark Pt/C (20 wt%) catalyst has been measured by linear sweep voltammetry

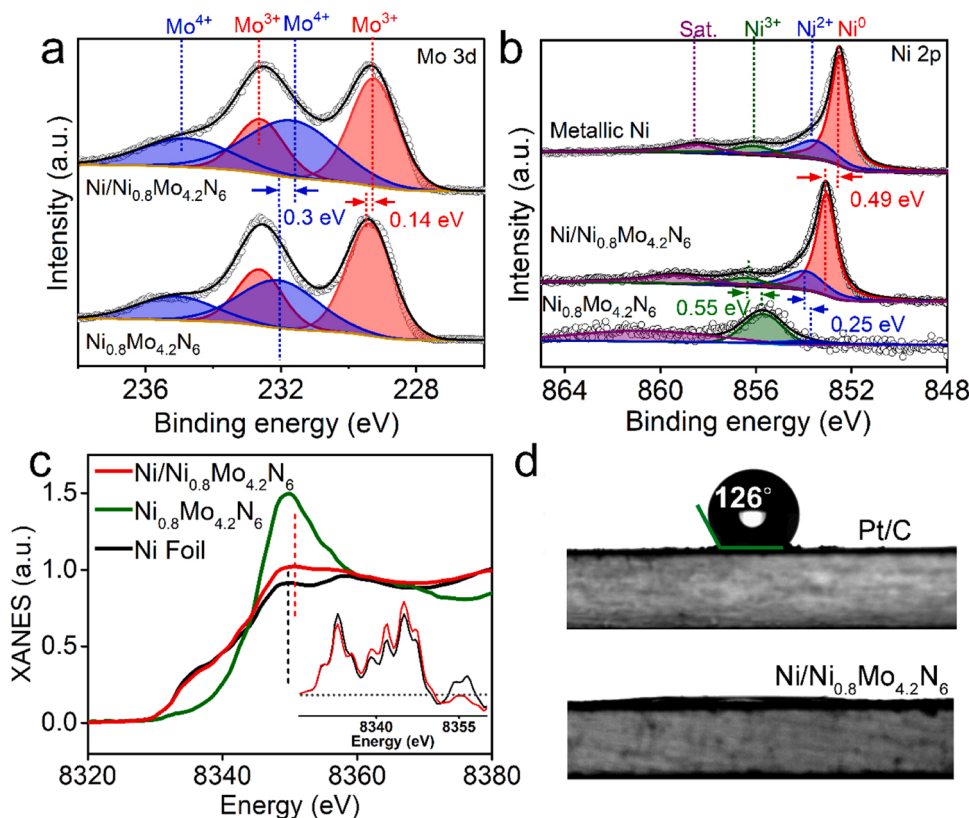


Fig. 2. (a) Mo 3d X-ray photoelectron spectroscopy (XPS) spectra of $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. (b) Ni 2p XPS spectra of metallic Ni, $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$. (c) Ni K-edge XANES spectra for the $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ in comparison with the reference compounds Ni foil (inset: expanded view of Ni K-edge pre-edge). (d) The contact angle measurement for hydrophilicity of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and Pt/C.

(LSV) methods. The results in Fig. 3a show the most positive potential for $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ suggesting the highest activity towards HER. At current density of 10 mA cm^{-2} , the overpotential of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is only 20 mV, much lower than those for either $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ or metallic Ni, and is even 21 mV lower than the benchmark Pt/C (20 mV). The $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ overpotential is found to be reproducible to $\pm 1.0 \text{ mV}$ (Fig. S8). Fig. S9 shows how the performance of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ for HER varies with the molar ratio of nickel, and the sample with 48.6 wt% cubic nickel is found to be optimal.

As an inherent property of an electrocatalyst, the Tafel slope is analyzed to reveal the HER mechanism. The Tafel slope for $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ was determined as 38 mV dec^{-1} by analyzing the minimum slope of the polarization curve (Fig. 3b). This suggests that the Volmer-Heyrovsky mechanism is responsible for the HER pathway [8, 40]. The step of electrochemical desorption ($\text{H}_{\text{ads}} + \text{H}_2\text{O} + \text{e}^- \rightarrow \text{H}_2 + \text{OH}^-$) is the rate-determining step (RDS) in this mechanism because of the high affinity of Ni for adsorbed H^+ . The Volmer reaction ($\text{H}_2\text{O} + \text{e}^- \rightarrow \text{H}_{\text{ads}} + \text{OH}^-$) plays a dominant role in activating water for the HER kinetics on both $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and metallic Ni. The Tafel slopes of $\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and metallic Ni are 180 and 123 mV dec^{-1} , respectively (Fig. S10). By extrapolating Tafel plots to zero overpotential, $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ has a J_0 of 2.89 mA cm^{-2} , five times greater than commercial Pt/C (0.49 mA cm^{-2}). The higher J_0 demonstrates that the intrinsic kinetics of the HER electron-transfer activity on $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is much faster than for Pt/C [41]. These results are further confirmed by electrochemical impedance spectroscopy (EIS) measurement. As shown in Fig. 3c, the charge transfer resistances for HER on $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ and Pt/C electrodes are 5.2 and 7.7Ω , respectively. This suggests that HER kinetics occur faster on $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ than on Pt/C. Additionally, the high (metallic) electrical conductivity of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ gives a resistance of 0.8Ω compared to 2.0Ω for Pt/C at the same mass loading. The superior conductivity of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is advantageous for

electrochemical catalysis to reduce the non-faradaic energy consumption [9]. The exceptional superhydrophilicity of $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ enables the aqueous electrolyte to effectively wet the electrocatalyst, thereby enhancing mass transfer in the HER process.

To determine the inherent electrocatalytic activity of different catalysts, the TOF of each Ni active site is computed by considering the total number of accessible Ni sites (Fig. S11). In Fig. 3d, at overpotentials of 50 and 100 mV, the respective TOF values for $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ are 0.29 and 0.76 s^{-1} , compared to 0.17 and 0.48 s^{-1} for the commercial Pt/C. Hence, in terms of overpotential, J_0 and TOF, $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ outperforms Pt/C as an efficient HER catalyst in alkaline solutions.

The practical applicability of an electrocatalyst heavily relies on its ability to maintain durability, which is a critical parameter that must be taken into consideration. A prolonged electrocatalytic test for hydrogen evolution reaction (HER) was conducted by cycling through four different current densities ($10\text{--}40 \text{ mA cm}^{-2}$) within a 24-hour interval. After 1000 h, the $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ displays a smaller potential decrease (26 mV) at 10 mA cm^{-2} current density than the commercial Pt/C (43 mV) in Fig. 3e. The HR-TEM images in Fig. S12 disclose no structural variations for $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ after the durability test, demonstrating structural robustness during alkaline HER electrocatalysis. The Faradic efficiency (FE) of the $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ is determined from the volume of hydrogen evolution (Fig. S13). Experimental results have demonstrated that the efficiency of H_2 volume production during electrolysis in a 1.0 M KOH solution is almost 100% ($>99.3\%$). This validates that the quantity of electricity passed is directly proportional to the amount of H_2 generated through the current flow [42–44].

Overall, the above results demonstrate that $\text{Ni}/\text{Ni}_{0.8}\text{Mo}_{4.2}\text{N}_6$ shows outstanding HER performance, superior to that of Pt/C and other reported state-of-the-art catalysts. Direct comparisons of catalyst overpotentials normalized by of electrode surface area, and of mass currents normalized by mass loading of catalyst are shown in Fig. 3f and g, with

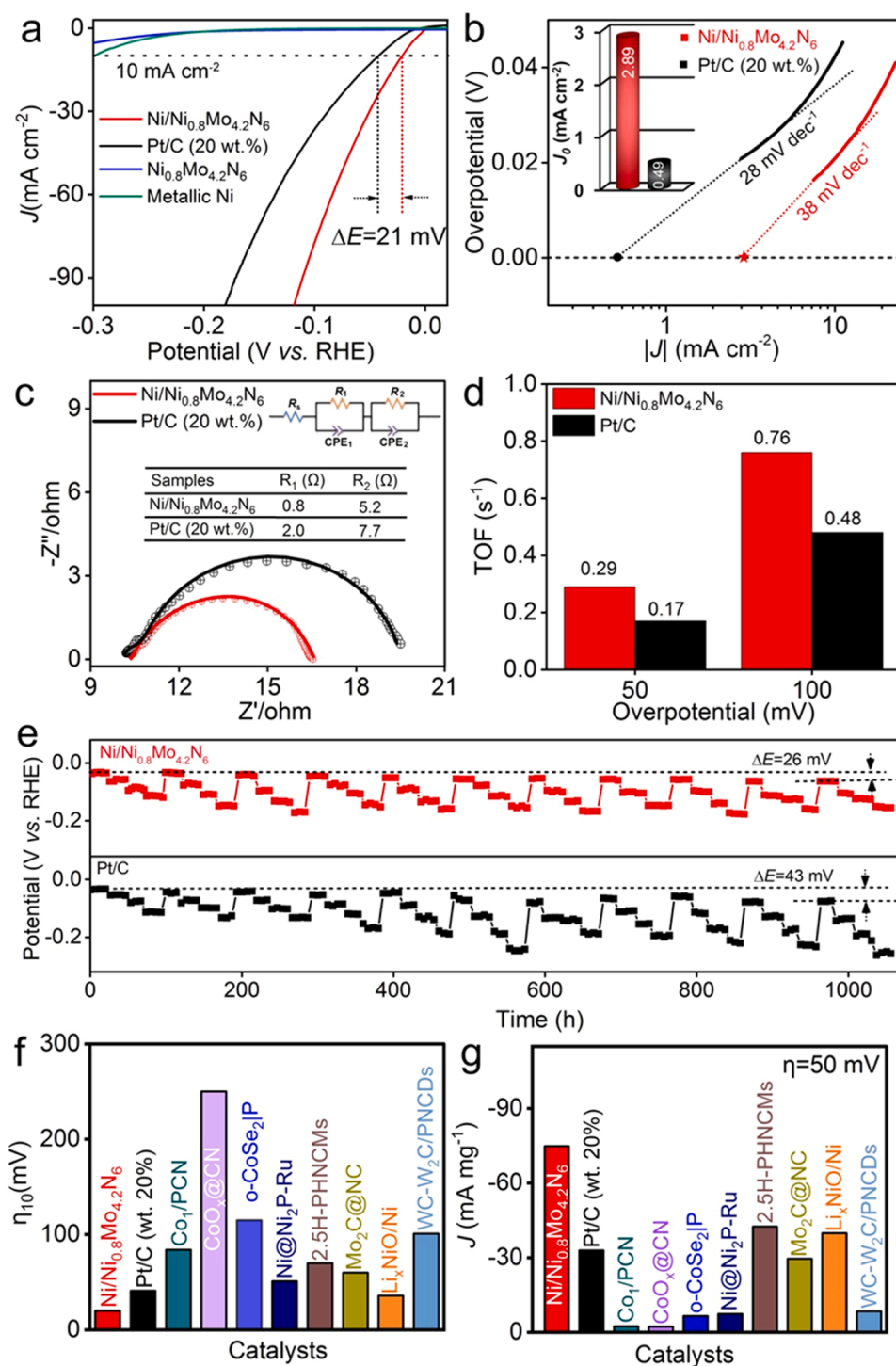


Fig. 3. (a) LSV polarization curves of the metallic Ni, Ni_{0.8}Mo_{4.2}N₆, Ni/Ni_{0.8}Mo_{4.2}N₆ and Pt/C (20 wt%) catalysts in 1.0 M KOH. (b) Tafel plots of the metallic Ni, Ni_{0.8}Mo_{4.2}N₆, Ni/Ni_{0.8}Mo_{4.2}N₆ and Pt/C catalysts. Insert shows the electrochemical J_0 results of the electrocatalysts are obtained by extrapolating the Tafel plots. (c) EIS curves of Ni/Ni_{0.8}Mo_{4.2}N₆ and Pt/C (20 wt%) electrocatalysts recorded during hydrogen evolution at $\eta = -100$ mV. (d) TOF of the Ni/Ni_{0.8}Mo_{4.2}N₆ and Pt/C toward HER in 1.0 KOH. (e) Cycling stability at multicurrent density of 10, 20, 30 and 40 mA cm⁻² without IR compensation. (f) The overpotential at 10 mA cm⁻² and (g) the mass current density at 50 mV overpotential for the reported materials.

further comparison data in Fig. S14 and Table S3.

To investigate the metal-support interactions of Ni and Ni_{0.8}Mo_{4.2}N₆, we analyzed the charge density distributions and conducted DFT calculations on the intermediate H* adsorption free energies (ΔG_{H^*}), which are detailed in the Supporting Information. The optimized structures of a Ni cluster, Ni_{0.8}Mo_{4.2}N₆ and Ni/Ni_{0.8}Mo_{4.2}N₆ are shown in Fig. S15. To investigate the impact of Ni_{0.8}Mo_{4.2}N₆ and Ni on the electronic structure, the density of states (DOS) was calculated and presented in Fig. 4b. All systems exhibited finite DOS values around the Fermi level, indicating their inherent metallic properties. The HOMO of Ni/Ni_{0.8}Mo_{4.2}N₆ located in close proximity to the Fermi level is closer than that of pure

metallic Ni. This suggests that Ni/Ni_{0.8}Mo_{4.2}N₆ possess a greater tendency to lose electrons, which weakens the H* adsorption and enhances the HER activity. These findings align with the outcomes from XPS and XANES investigations. Ni/Ni_{0.8}Mo_{4.2}N₆ shows excellent performance, which is attributed to the low dissociation barrier of water for Ni-based catalyst materials. As shown in Fig. 4c, the energy barrier for breaking the OH-H bond in water is 1.19 eV on Pt (111) surface and such a high energy barrier clearly hinders the dissociation of water to H_{ads} [29]. Strikingly, the dissociation barrier of water is 0.52 eV on Ni/Ni_{0.8}Mo_{4.2}N₆ surface. Thus, the Ni/Ni_{0.8}Mo_{4.2}N₆ surface can promote water dissociation substantially and increase the rate of H_{ads} formation. The

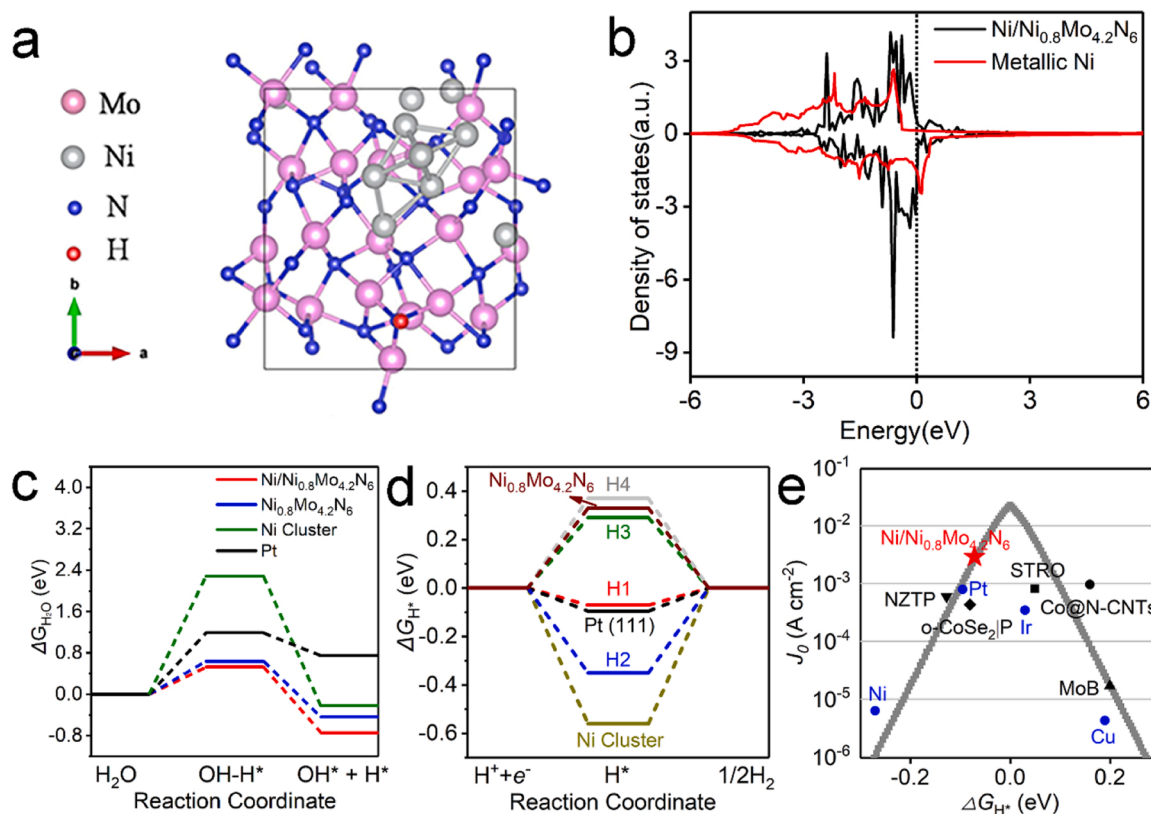


Fig. 4. (a) Optimized structures of Ni/Ni_{0.8}Mo_{4.2}N₆. (b) Calculated electronic densities of states metallic Ni and Ni/Ni_{0.8}Mo_{4.2}N₆. (c) Reaction energy diagram of water dissociation on the Pt (111) surface (black), Ni Cluster (green), Ni_{0.8}Mo_{4.2}N₆ (blue) and Ni/Ni_{0.8}Mo_{4.2}N₆ surface (red). The initial state (H₂O), the transition state (TS) and the final state (H* + OH*) are indicated in the diagram with the corresponding energy barrier on the two surfaces. (d) Calculated free-energy diagram of HER over metallic Ni cluster, Pt (111) and Ni/Ni_{0.8}Mo_{4.2}N₆ at equilibrium potential. (e) Volcano plot of experimentally measured J_0 as a function of the DFT-calculated Gibbs free energy of adsorbed atomic hydrogen. Detailed cited references can be found in Table S3 in the Supporting Information.

Sabatier principle suggests that a site with a moderate Gibbs free energy value ($\Delta G_{H^*} = 0$ eV) for atomic hydrogen adsorption is optimal for the HER [15]. As shown in Fig. 4d, the calculated ΔG_{H^*} on the site of H1 in Ni/Ni_{0.8}Mo_{4.2}N₆ is -0.07 eV, more close to zero than for Pt. And much larger than the Ni cluster (-0.55 eV) and the Ni_{0.8}Mo_{4.2}N₆ (0.33 eV) (Fig. S16). This reveals more favorable H* desorption kinetics on Ni/Ni_{0.8}Mo_{4.2}N₆ during the HER process. Ni/Ni_{0.8}Mo_{4.2}N₆ is located near the top of the J_0 vs. ΔG_{H^*} volcano diagram in Fig. 4e, closely resembling state-of-the-art hydrogen electrocatalysts (Table S3). This confirms the outstanding performance of Ni/Ni_{0.8}Mo_{4.2}N₆.

4. Conclusions

In conclusion, Ni/Ni_{0.8}Mo_{4.2}N₆ demonstrates improved catalytic kinetics and offers high activity for HER owing to the strong interaction between the Ni nanoparticle and Ni_{0.8}Mo_{4.2}N₆. The combination of surface superhydrophilicity and electron withdrawal reduce the desorption energy of hydrogen on nickel, and allows the Ni/Ni_{0.8}Mo_{4.2}N₆ electrocatalyst to exhibit impressive ultra-high HER activity with excellent stability over 1000 h in alkaline media. Application of Ni/Ni_{0.8}Mo_{4.2}N₆ in water splitting, chlorine alkali industry and related electrochemical energy conversion and storage devices is anticipated. Our strategy of engineering catalyst particle local electronic structure through substitution of the same element within the support may also lead to discovery of other outstanding materials for the HER and related electrochemical processes.

CRediT authorship contribution statement

Shuqin Liang: Investigation, Data curation, Formal analysis,

Writing – original draft. **Huashuai Hu:** Investigation, Formal analysis, Methodology, Writing – original draft. **Jue Liu:** Formal analysis, Writing – original draft. **Hangjia Shen:** Investigation, Methodology. **Qiao Li:** Investigation, Methodology. **Nianxiang Qiu:** Investigation, Methodology. **HaiChuan Guo:** Software, Formal analysis. **Xuyun Guo:** Investigation. **Shiyu Du:** Software, Formal analysis. **Ye Zhu:** Conceptualization. **Jian Liu:** Supervision. **J. Paul Attfield:** Conceptualization, Writing – review & editing. **Minghui Yang:** Supervision, Conceptualization, Writing – review & editing, Funding acquisition.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

Data availability

No data was used for the research described in the article.

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Appendix A. Supplementary material

Supplementary data associated with this article can be found in the online version at doi:10.1016/j.apcatb.2023.123008.

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